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ULTRAVOILET-VISIBLE SPECTROSCOPY: A REVIEW

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ABSTRACT
Ultraviolet and visible (UV-Vis) absorption spectroscopy is the measurement of the attenuation of a beam of light after it passes through a sample or after reflection from a sample surface. This article uses the term UV-Vis spectroscopy to include a variety of absorption, transmittance, and reflectance measurements in the ultraviolet (UV). Visible and near-infrared (NIR) spectral regions. These measurements can be at a single wavelength or over an extended spectral range. This article provides an overview of the technique and does not attempt to provide a comprehensive review of the many applications of UV-Vis spectroscopy in materials research. In this regard, many of the references were chosen to illustrate the diversity of applications rather than to comprehensively survey the uses of UV-Vis spectroscopy. Rapid and easy analytical methods are needed due to increasing number of multicomponent formulation bio therapeutic product and sample of complex matrix Number of ultraviolet spectrophotometer method use for these purposes. Ultraviolet and visible absorption Spectroscopy is the measurements of the attenuation of a beam of light after reflection from a sample surface.

KEYWORDS: Ultraviolet-Visible-near Infrared (UV-VIS-NIR), Ultraviolet-Visible (UV-VIS), Light emitting diode (LED). Transmittance.

INTRODUCTION
What is UV-Vis spectroscopy? UV-Vis spectroscopy is an analytical technique that measures the amount of discrete wavelengths of UV or visible light that are absorbed by or transmitted through a sample in comparison to a reference or blank sample. Ultraviolet-visible (UV-visible) spectrophotometry is primarily a quantitative analytical technique concerned with the absorption of near-UV (180-390 nm) or visible (390-780 nm)radiation by chemical species in solution. Spectroscopy in the ultraviolet (UV), visible (Vis) and near infrared (NIR)region of the electromagnetic spectrum.

The Principle of UV-Visible Spectroscopy is based on the absorption of ultraviolet light or visible light by chemical compounds, which results in the production of distinct spectra. Spectroscopy is based on the interaction between light and mutter. • Spectroscopic investigations of solutions, gas phase crystals usually take place in transmission, but it is very difficult to obtain transparent films of powders (e.g., heterogeneous catalysts), making transmission experiments almost impossible. Alternatively, diffuse reflected light can be collected and this technique has been named diffuse reflectance spectroscopy (DRS). One of the advantages of DRS is that the obtained information is directly chemical in nature since outer shell electrons of the ions are probed. This provides information about state and coordination environment of ions in catalytic solids. The same holds for the nature species and different hydrocarbon species can be investigated. Furthermore, DRS is quantitative and can be in-situ conditions. The main disadvantage of the technique is that DRS spectra are complex, and usually encompass several broad and overlapping bands.

In order to spectral analysis, techniques need to. This is especially important for in-situ time-re-solved DRS studies because of the extensive to be handled. This chapter starts with a short overview of the principles of DRS. Theoretical as well as practical aspects will be discussed. The next section focuses on three examples in order to illustrate the potential and limitations of in situ NIR spectroscopy.UV visible spectroscopy is the classical and the most reliable technique for qualitative and quantitative analysis of organic compounds visible spectroscopy is an analytical technique that measures the amount of discrete wavelengths of UV or visible light that are absorbed by or transmitted through a sample in comparison to a reference or blank sample.

PRINCIPLE	SUBTYPES	EXAMPLE
Study is based upon atomic or molecular level	Atomic Spectroscopy	Flame photometry, Atomic Absorption Spectroscopy
	Molecular Spectroscopy	UVspectroscopy, Colorimetry
Study is based upon absorption or emission of EMR	Absorption Spectroscopy	Infra-Red and NMR Spectroscopy
	Emission Spectroscopy	Fluorimetry, Flame photometry
Study is based upon electronic or magnetic levels	Electronic Spectroscopy	UV spectroscopy, Colorimetry, Fluorimetry
	Magnetic Spectroscopy	NMR and ESR spectroscopy

The energy of a molecule can be due to electronic, vibrational or rotational energy. They are in the following ratio:
Rotational energy: Vibrational energy: electronic Energy=1:100:10,000.

PRINCIPLE OF UV VISIBLE SPECTROSCOPY
When radiation in the wavelength range 10 to 780 nm passes through the solution of the compound electron get excited from lower energy level to higher energy level during this process it absorption some energy to the solution. The difference between the energies of ground state and higher excited state is called as delta E and this energy is equal to the amount of UV radiation absorb by the molecule. UV visible spectroscopy is the most important techniques for qualitative or quantitative analysis it is also called as electronic spectroscopy. The Principle of UV-Visible Spectroscopy is based on the absorption of ultraviolet light or visible light by chemical compounds, which results in the production of distinct spectra. Spectroscopy is based on the interaction between light and matter. When the matter absorbs the light, it undergoes excitation and de-excitation. Resulting in the production of a spectrum. When matter absorbs ultraviolet radiation, the electrons present in it undergo excitation. This causes them to jump from a ground state (an energy state with a relatively small amount of energy associated with it) to an excited state (an energy state with a relatively large amount of energy associated with it). It is important to note that the difference in the energies of the ground state and the excited state of the electron is always equal to the amount of ultraviolet radiation or visible radiation absorbed by it. Spectrophotometry is a procedure for determining how much light is reflected by a chemical material by measuring the strength of light as a light beam travels through the sample solution. The fundamental theory is that light is absorbed or emitted over a certain wavelength spectrum by each compound. Spectroscopy is known as the measurement and interpretation of electromagnetic radiation emitted or absorbed, when ions, atoms or molecules of a sample move from one energy state to another energy state. Ultraviolet-Visible spectroscopy UV-Vis spectroscopy is a type of absorption spectroscopy, which utilizes the radiation in the UV range and adjacent visible range of the electromagnetic radiation spectrum. The absorption wavelength mainly ranges from 100-700 nm.

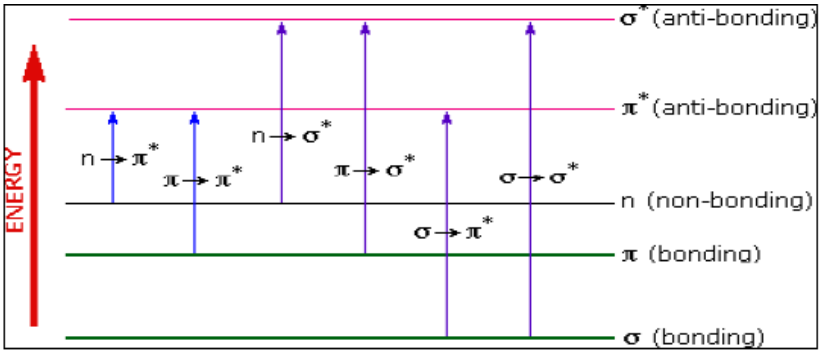
ELECTRONIC TRANSITION
The absorption of ultraviolet or visible radiation generally results from excitation of bonding electron as a consequence, the wavelength of absorption peaks can be correlated withthe types of bonds that exist in the species under study UV visible spectroscopy involve transition of electrons to from lower energy level to higher energy level. When any molecule absorb UV radiation their outermost electrons absorb the radiation and get excited to higher energy level.

THREE TYPES OF ELECTRON
1. π ELECTRON
2. SIGMA ELECTRON
3. N-ELECTRON

1. π ELECTRON
Compounds having double and triple bonds undergo this type of transition electron. Example -- Benzene. The bonding molecular orbital component of a pi bond. The orbital of ethylene's carbon- carbon pi bond has two orbital lobes, one above the plane of the atoms, and another below the plane. This is a bonding molecular orbital. The plane containing the atoms is also the pi orbital's one node.

2. SIGMA ELECTRON
These electrons present in all single bonded compounds. Example-methane, propane. Sigma bonds are the strongest type of covalent chemical bond. They are formed head-on overlapping between atomic orbitals. Sigma bonding is most simply defined for diatomic molecules using the language and tools of symmetry groups.

3. N ELECTRON (non bonding electron)
These electrons are not involve in bond formation. They are absolutely free or easily available. Example nitrogen oxygen sulfur. Non-bonding orbitals are often designated by the letter n in molecular orbital diagrams and electron transition notations.



FOUR TYPES OF ELECTRON TRANSITION
1) SIGMA TO SIGMA STAR
2) N TO SIGMA STAR
3) π TO π STAR
4) N TO π STAR

1. SIGMA TO SIGMA STAR
Saturated compounds undergo this type of transition. Example - methane, propane, ethane. Single bonded compounds. An electronic in a bonding Sigma orbital of a molecule is excited to the corresponding anti- bonding orbital by the absorption of radiation. In ethane, electrons of C-C bond appear to be involved. Because the strength of the C-C bond is less than that of the C-H bond, less energy is required for excitation thus the absorption peaks occurs at a longer wavelengths.

2. N- TO SIGMA STAR
This type of transition involves compound having at least one hetero atom with lone pair of electron. It is less than sigma to sigma star N to sigma star transition (n→→ 0*) involves saturated compounds with one hetero atom like oxygen, nitrogen, fluorine, chlorine, etc. Normally, saturated halides, alcohols, ethers, aldehyde, ketones, and amines participate in this type of transition. Absorption maxima for the formation of the n. sigma star state tend to shift to shorter wavelengths in the presence of polar solvents such as water or ethanol.

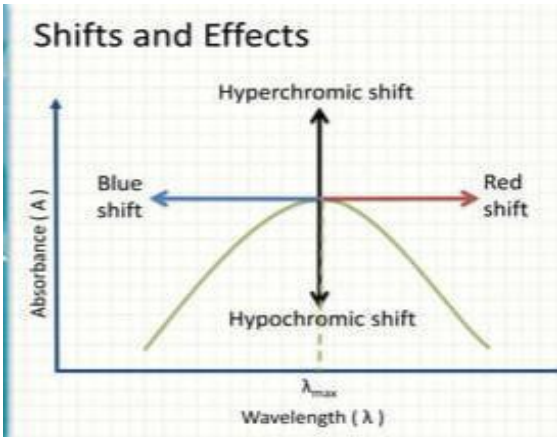
3. N- TO π STAR
Compounds having double and triple bonds undergo this type of transition. All aromatic compounds. Example - Benzene. These transitions involve moving an electron from a bonding a orbital to an ant bonding ** Orbital. They tend to have molar absorptivity on the order of 10,000 and undergo a red shift with solvent interactions (a shift to lower energy and longer wavelengths).

4. N TO π STAR
It involves very less energy as compare to other transition. Compounds having double and triple bonds at least one hetero atom undergo this type of transition. The "n" electrons (or the nonbonding electrons) are the ones located on the oxygen of the carbonyl group of tetraphenyclopentadienone.

CHROMOPHORE AND AUXOCHROME
CHROMOPHORE: This term used for any group which gives colour to the compound any group which absorb UV and visible radiation and shows pic in this spectrum is called as chromophore Both organic and inorganic molecules may exhibit absorption and emission of UV-VIS radiation. Molecular groups that absorb visible or UV light are called chromophores. A chromophore group is a functional group, not conjugated with another group, which exhibit a characteristic absorption spectrum in the ultraviolet or visible region. The chromophore is a region in the molecule where the energy difference between two separate molecular orbitals falls within the range of the visible spectrum. Visible light that hits the chromophore can thus be absorbed by exciting an electron from its ground state into an excited state.

A chromophore is the section of a molecule that causes us to see color. The chromophore portion of the molecule will have alternating double bonds, or conjugated double bonds. For example, beta-carotene, the molecule responsible for the color in carrots, has many double bonds.

AUXOCHROMES: Any group which dose not act as chromophore but it's presence shift the absorption maxima towards higher wavelengths. The colour of molecules may be interested by the groups called as auxochromes which generally do not absorb significantly in the 200-800 nm region, but will affect the spectrum of the chromophore to which it is attached.The most important auxochromes groups are OH, NH2, CH3 and NO2 and their properties are acidic or basic An auxochrome is a functional group of atoms with one or more lone pairs of electrons when attached to a chromophore, alters both the wavelength and intensity of absorption.



FOUR SHIFTING OF CHROMOPHORE
1) BATHOCHROMIC SHIFT
2) HYPSOCHROMIC SHIFT
3) HYPERCHROMIC SHIF
4) HYPOCHROMIC SHIFT

1. BATHOCHROMIC SHIFT (RED SHIFT)
Shifting of absorption maxima towards higher wavelengths due to addition of auxochromes groups is called as BATHOCHROMIC shift and also called as red shift. Bathochromic shift is a change of spectral band position in the absorption, reflectance. Transmittance or emission spectrum of a molecule to a longer wavelength. Because the red color in the visible spectrum has a longer wavelength than most other colors, the effect is also commonly called a red shift. Due to the presence of an auxochrome, or solvent effect is called a bathochromic shift or red shift. For example, benzene shows Amax 256 nm and aniline shows max 280nm.

2. HYPSOCHROMIC SHIFT (BLUE SHIFT)
• Shifting of absorption maxima towards lower wavelengths due to removal of auxo group is called as hypsochromic shift and also called as blueshift. Hypochromic shift is a change of spectral band position in the absorption, reflectance. Transmittance or emission spectrum of a molecule to a shorter wavelength. Because the blue color in the visible spectrum has a shorter wavelength than most other colors, this effect is also commonly called a blue shift.

• A hypsochromic shift is the shift of a peak or signal to shorter wavelength (higher energy). Also called a blue shift. For an absorption peak starting at Amax = 550 nm, a shift to higher wavelength such as 650 nm is bathochromic, whereas a shift to lower wavelength such as 450 nm is hypsochromic

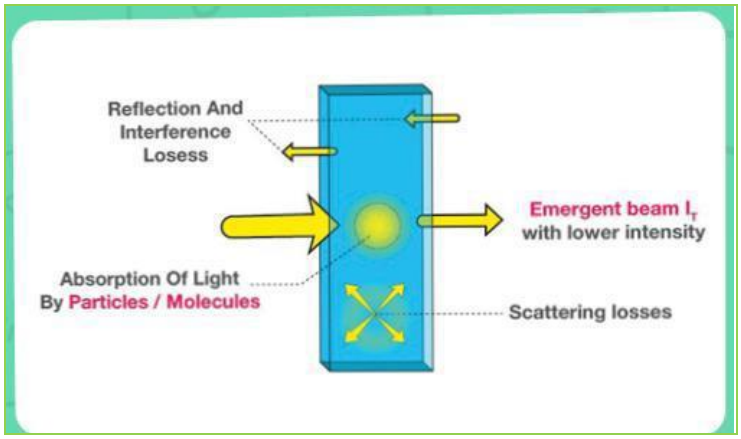
3. HYPERCHROMIC SHIFT
Shifting in the intensity of absorption maximum towards higher values of absorbance is called as hyperchromic shift. An increase in the absorption of ultraviolet light by a solution of DNA as these molecules are subjected to heat, alkaline conditions, etc. The shift is caused by the disruption of the hydrogen bonds of each DNA duplex to yield single-stranded structures. • The phenomenon of UV absorbance increasing as DNA is denatured is known as the hyperchromic shift. The purine and pyrimidine bases in DNA strongly absorb ultraviolet light. Double-stranded DNA absorbs less strongly than denatured DNA due to the stacking interactions between the bases.

4. HYPOCHROMIC SHIFT
Shift in the intensity of absorption maxima towards lower absorbance is called as hypochromic shift.The Hypochromic Effect describes the decrease in the absorbance of ultraviolet light in a double stranded DNA compared to its single stranded counterpart. Compared to a single stranded DNA, a double stranded DNA consists of stacked bases that contribute to the stability and the hypochromicity of the DNA.

BEER-LAMBERT LAW
Beers Lamberts law states that the absorption 'A' of a substance in solution is directly proportional to the concentration of the solution whereas Lamberts law states that each layer of equal thickness of an absorbing medium absorbs an equal fraction of the radiant energy.

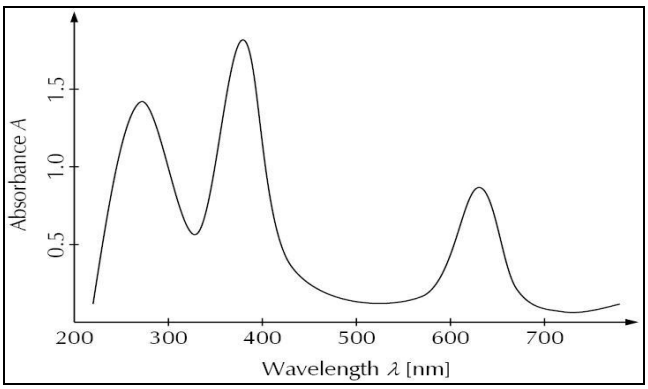
a) When passing through a transparent cuvette filled with sample solution, the light intensity is reduced proportional to the sample solution concentration. In other words, a higher concentrated sample solution will absorb more light. In addition the reduction the light intensity is also proportional to the length of the cuvette, a longer cuvette will be to a higher absorption of light.

b) The Beer-Lambert law states that there is a linear relationship between the concentration and the absorbance of the solution, which enables the concentration of a solution to be calculated by measuring its absorbance.



FORMULA
A=a b e Where,
Where,
A= Absorbance
a = Molecular absorbing coefficient of the species.
b= Absorbing layer (path length) and
C = Molecular concentration of the absorbing species.

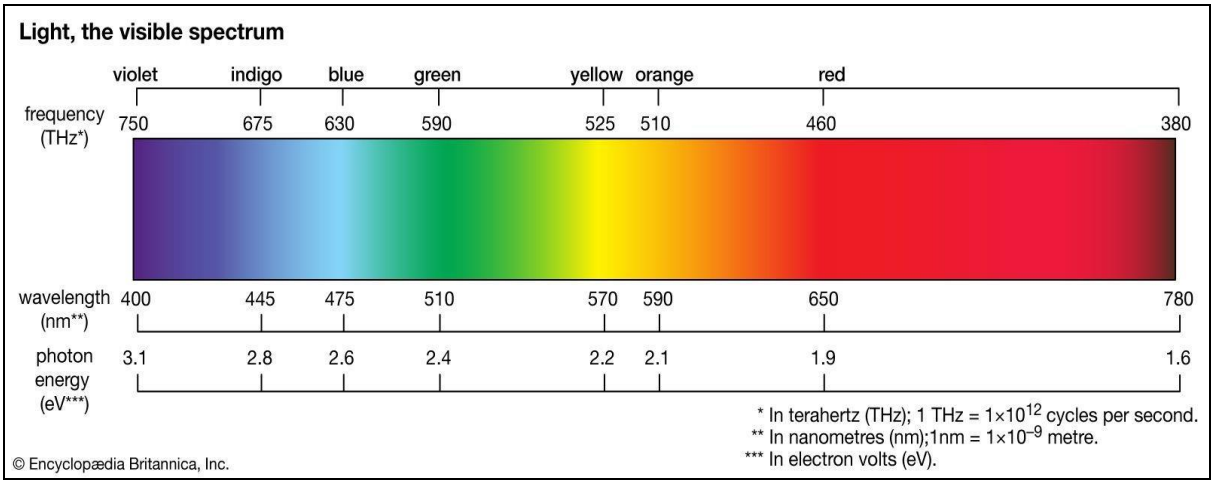
6) SPECTROSCOPY
When an electromagnetic radiation is incident on a matter, phenomena like reflection, transmission, absorption are occurring. Spectroscopy is the study of interaction of electromagnetic radiation with matter based on the Bohr-Einstein Frequency relationship E- hv here h is the proportionality constant called planks constant and V is frequency. Measurements of radiation intensity as a function of wavelengths are described by spectroscopy.



Interaction between electromagnetic radiation and matter as a function of the wavelength or frequency of the radiation. Spectroscopy is used in physical and analytical chemistry to detect, determine, or quantify the molecular and/or structural composition sample. Each type of molecule and atom will reflect, absorb, or emit electromagnetic radiation in its own characteristic way.

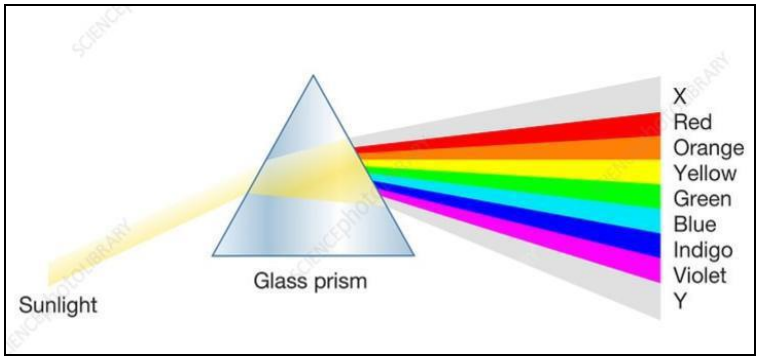
Spectroscopy is the study of the absorption and emission of light and other radiation by matter. It involves the splitting of light (or more precisely electromagnetic radiation) into its constituent wavelengths (a spectrum), which is done in much the same way as a prism splits light into a rainbow of colours.

SPECTRUM: The spectrum is formed by electromagnetic waves and the wavelength is varies. Plural spectra spectra or spectrums. a continuum of color formed when a beam of white light is dispersed (as by passage through a prism) so that its component wavelengths are arranged in order. Autism is known as a "spectrum" disorder because there is wide variation in the type and severity of symptoms people experience.



When a narrow beam of light is allowed to pass through a prism, grating it is dispersed into seven colors from red, violets and band is called spectrum. When white light is passed through a glass prism it splits into its spectrum of colours (in order violet, indigo, blue, green, yellow, orange and red) and this process of white light splitting intoits constituent colours is termed as dispersion. A word was first used scientifically in optics to describe the rainbow of colours invisible light after passing through a prism as scientific understanding of light advanced it came to apply to the entire electromagnetic spectrum.

Glass prism dispersion



RAINBOW

Colors A rainbow shows up as a spectrum of light: a band of familiar colors that include red, orange, yellow, green, blue, and violet. The name "Roy G. Biv" is an easy way to remember the colors of the rainbow, and the order in which they appear: red, orange, yellow, green, blue, indigo, and violet.

This is the order of the wavelengths of visible light starting with red (the longest) and ending with violet (the shortest). It is also the order that colors appear in a rainbow!

Rainbows appear when sunlight shines through water droplets suspended in the atmosphere.

This is the order of the wavelengths of visible light starting with red (the longest) and ending with violet (the shortest). It is also the order that colors appear in a rainbow! Rainbows appear when sunlight shines through water droplets suspended in the atmosphere.

The only difference between UV and IR light versus visible light is the wavelength.

UV-VISIBLE SPECTROSCOPY

Ultraviolet visible spectrum - can be generated when ultraviolet light and visible light (200-900) are absorbed by materials. The spectrum can be used to analyze the composition and the structure of the material. For a particular wavelength in the ultraviolet visible ranges, the absorption degree is proportional to the components of the material.

UV spectroscopy or UV-visible spectrophotometry refers to absorption spectroscopy or reflectance spectroscopy in part of the ultraviolet and the full, adjacent visible regions of the electromagnetic spectrum.

Ultraviolet-visible (UV/Vis) spectroscopy is based on the absorption of the electromagnetic radiation in UV/Vis region, with the wavelength ranges of 200-400 nm, called 'ultraviolet spectroscopy, and 400-800 nm, called 'visible spectroscopy.

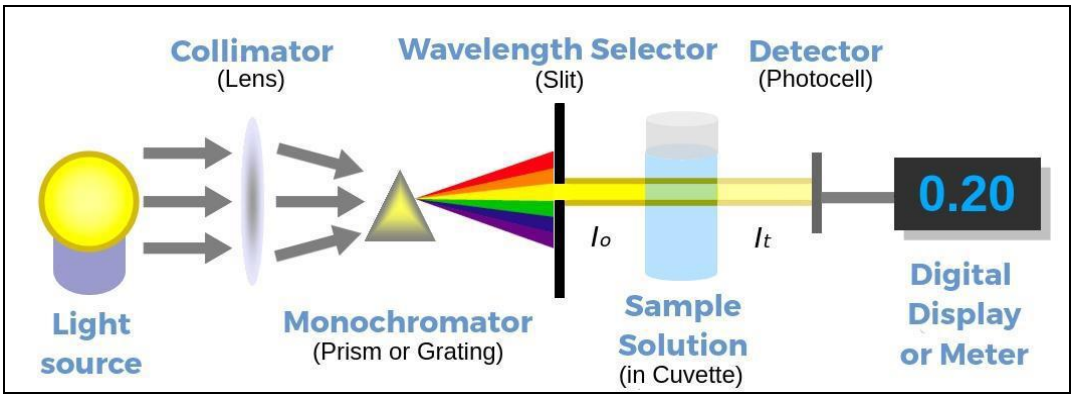
UV-Vis spectroscopy is an analytical technique that measures the amount of discrete wavelengths of UV or visible light that are absorbed by or transmitted through a sample in comparison to a reference or blank sample. This property is influenced by the sample composition, potentially providing information on what is in the sample and at what concentration.

The Principle of UV-Visible Spectroscopy is based on the absorption of ultraviolet light or visible light by chemical compounds, which results in the production of distinct spectra. Spectroscopy is based on the interaction between light and matter. Ultraviolet-visible (UV-Vis) spectroscopy is a widely used technique in many areas of science ranging from bacterial culturing, drug identification and nucleic acid purity checks and quantitation, to quality control in the beverage industry and chemical research.

UV spectroscopy or UV-visible spectrophotometry (UV-Vis or UV/Vis) refers to absorption spectroscopy or reflectance spectroscopy in part of the ultraviolet and the full, adjacent visible regions of the electromagnetic spectrum.

INSTRUMENTATION

1) SOURCES OF LIGHT
2) MONOCHROMATOR
3) SAMPLE SOLUTION IN CUVETTE
4) PHOTO DETECTOR
5) READOUT DEVICES



COMPONENTS OF SPECTROMETER

A spectrophotometer is an analytical instrument used for the objective calculation of visible light. UV light, or infrared light emission or reflection. Spectrophotometers measure intensity as a function of the wavelength of the light source.

A spectrophotometer measures the number of photons emitted to estimate the intensity of light spectra absorbed and transmitted by a sample. This provides information on the amount of a compound in the sample.

Spectrophotometry is a method to measure how much a chemical substance absorbs light by measuring the intensity of light as a beam of light passes through sample solution. The basic principle is that each compound absorbs or transmits light over a certain range of wavelength. This measurement can also be used to measure the amount of a known chemical substance.

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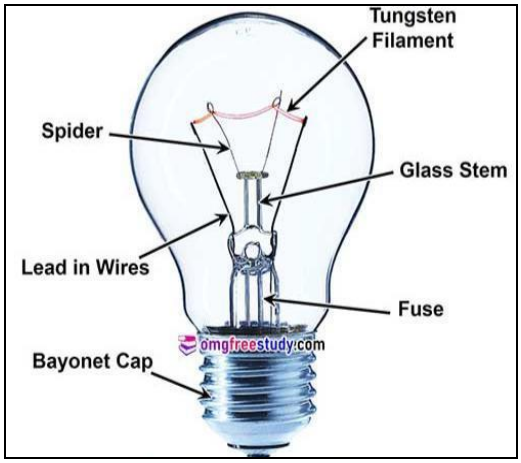
1. SOURCES OF LIGHT

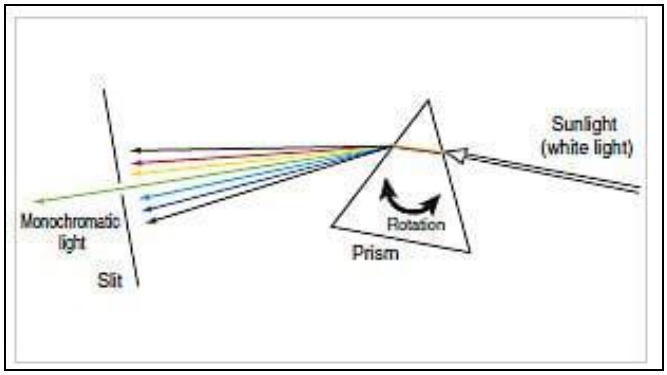
Radiation sources should be economic. It should be long lasting show should be stable over time. Example - Tungsten lamp. hydrogen discharge lamp, Deuterium lamp, xenon dischargelamp, mercury are Part of the UV and visible radiation source is Tungsten lamp. A light source is anything that makes light, whether natural and artificial. Natural light sources include the Sun and stars. Artificial light sources include lamp posts and televisions.



Tungsten filament incandescent lamps, particularly tungsten halogen lamps, are often used in illumination systems UV radiation sources is Deuterium or hydrogen lamp. Range of wavelengths 200-400 nm. Deuteriums are lamp is a low-pressure gas-discharge light source often used in spectroscopy when a continuous spectrum in the ultraviolet region is needed. Plasma "are" or discharge lamps using hydrogen are notable for their high output in the ultraviolet, with comparatively little output in the visible and infrared.

2. MONOCHROMATOR

It is a device that breaks the polychromatic radiation into comments wavelengths. It converts polychromatic light into monochromatic light. A monochromatic is an optical device that transmits a mechanically selectable narrow band of wavelengths of light or other radiation chosen from a wider range of wavelengths available atthe input.



MONOCHROMATOR

A monochromatic is an optical instrument which measures the light spectrum. Light is focused in the input slit and diffracted by a grating. In this way, only one color is transmitted through the output slit at a given time. Spectra are then recorded wavelength by wavelength, rotating the grating. There are modified types of monochromatic, for example the Fastie-Ebert monochromatic wit a common collimator/refocusing mirror, and devices with two gratings for better resolution. The quality of the diffraction grating can be important for the performance: Its diffraction efficiency determines the power losses.

- Monochromatic is a mechanism that emits monochromatic light from a light source. A dispersive element, generally a prism or diffraction grating, is used to create the monochromatic light.
- There are two types of monochromatic: prisms and grating systems.

The most common materials for laboratory X-ray monochromatic are pyrolytic graphite for broad band use and silicon, germanium, or quartz for narrow band use.

THE MONOCHROMATOR UNIT CONSISTS OF

- **ENTRANCE SLIT:** Definition narrow beam of radiation from source provide a narrow optical image of the radiation source. A thin slit in an opaque screen by which light enters a spectrometer. The spectrum thus formed is the image of this slit in each wavelength of light present.
- **COLLIMATING MIRROR:** A collimator is a device which narrows a beam of particles or waves. To narrow can mean either to cause the directions of motion to become more aligned in a specific direction, or to cause the spatial cross section of the beam to become smaller.

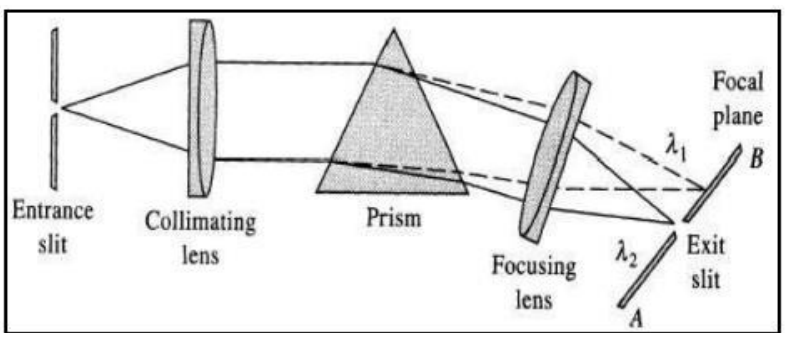
- In optics, a collimator may consist of a curved mirror or lens with some type of light source and/or an image at its focus.

DIFFRACTION GRATING OR PRISM

- Make of quartz disperses the light into specific wavelengths. The prism achieves dispersion due to the difference in the material refractive index according to the wavelength. However, the diffraction grating uses the difference in diffraction direction for each wavelength due to interference.

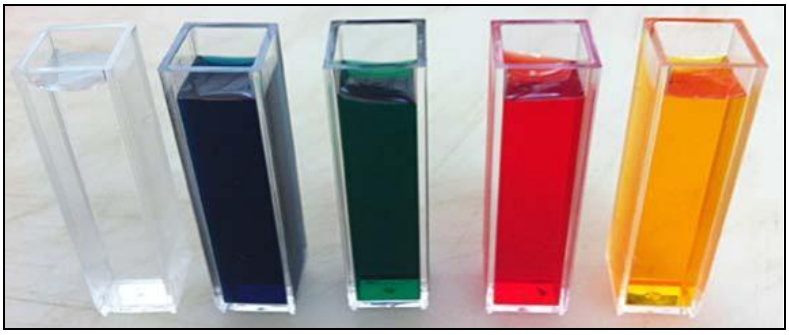
FOCUSING MIRROR

- Capture the dispersed light and sharpens the same to the sample via exit slit.
- A focusing mirror re-forms the image of the entrance slit and focuses it onto the exit slit.
- Focusing mirrors (concave mirrors) are characterized by one concave surface with a high reflection coating. The reflective coating allows focusing of a light beam.



3. SAMPLE SOLUTION IN CUVETTE

- Liquid sample is usually contained in a cell called as cuvette. Fingerprints are droplets of water disrupt light rays during measurements. • Cuvette from Quartz can be used in UV as well as in visible spectroscopy. Cuvette from glass is suitable for visible but not for UV spectroscopy because it absorb UVradiation.



Sample Solution In Cuvette.

Often the sample is a solution, with the substance of interest dissolved within. The sample is placed in a cuvette and the cuvette is placed in a spectrophotometer for testing. The cuvette can be made of any material that is transparent in the range of wavelengths used in the test.

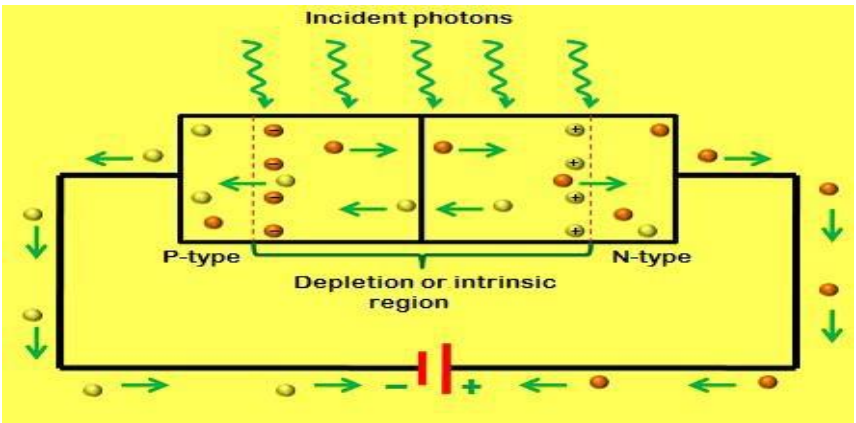
A cuvette is a type of sample holder for liquid samples. Often, they are made of plastic, borosilicate glass, or quartz. Stellar Net offers glass cuvettes for experiments in the visible or NIR ranges and quartz cuvettes for experiments in the UV range. Cuvettes also come with two or four polished sides.

When you rinse the cuvettes with water, the water dilutes the sample (concentration changes). Therefore, you must rinse the cuvettes with the sample (rather than water) so you can avoid changing the concentration of your sample.

4. PHOTO DETECTOR

A photo detector is a semiconductor device which converts light energy to electrical energy. It consists of a sample P-N junction diode and is designed to work in reverse biased condition. The photons approaching the diode are absorbed by the photodiode and current is generated.

Photo detectors, also called photo sensors, are sensors of light or other electromagnetic radiation. There is a wide variety of photo detectors which may be classified by mechanism of detection, such as photoelectric or photochemical effects, or by various performance metrics, such as spectral response.



Photodiode

Photo detectors are sensors that can convert the photon energy of light into electrical signal. The photo detector is also called an optical receiver. It converts the variation in optical power into a corresponding variation in the electric current. As compare to the optical transmitter the design of optical receiver is more complicated just because the receiver must detect weak, distorted signals and then make decisions on what type of data was sent based on an amplified version of this totally distorted signal.

TYPES OF DETECTOR

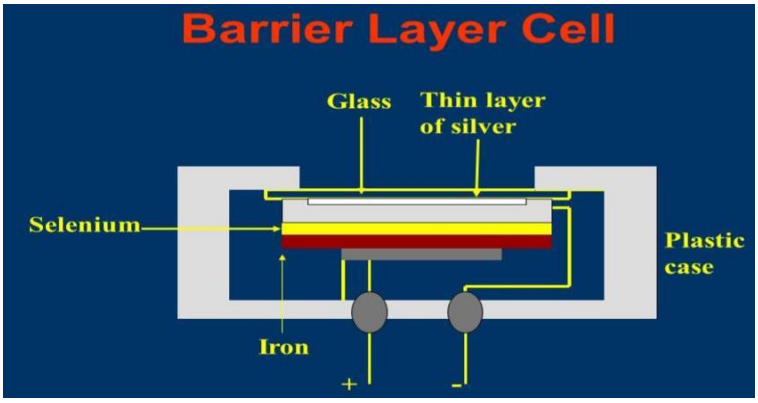
1. BARRIER LAYER CELL DETECTOR
2. PHOTO TUBE DETECTOR
3. PHOTO MULTIPLIER TUBE DETECTOR
4. SILICON PHOTODIODE DETECTOR

1. BARRIER LAYER CELL DETECTOR

Barrier layer cell detector is also called as photo volatile cell. It consists of semiconductor (selenium) which is deposited on strong base ion A very thin layer of silver of gold is placed over the surface of semiconductor.

• To act as collector electro radiation when falls on the surface electron on produce. Electron is produce and their it is converted into electric current. To dose not required power supply.

A photoelectric detector which is made of iron coated with a semiconductor film when light from 250-750nm hits this cell, you get a current; this is a cell which is mainly good for intense light sources, because there is not a huge signal enhancement. Also known as a self-generating barrier layer cell. A photoelectric detector that converts radiant flux directly into electrical current. Generally, it consists of a thin silver film on a semiconductor layer deposited on an iron substrate.



5. READOUT DEVICES

The signal from detector is received by recording system. Recording is done by recording pen. The earliest instruments were simple and directly connected the amplifier detector signal to a chart recorder. Nowadays, all experimental settings are controlled by a computer and detector signal are digitized processed and stored. The device used to accept the signal transmitted from the analyzer and display it for use in process operation or other decision making. Displayed as a measured property or concentration in the accepted units of that property or concentration.

Several types of readout devices are used in modern instruments. These devices include Digital Meters, Recorders, Cathode-Ray Tubes, LCD panels, and Computer Displays. This image shows an example of what readout will look like from the signal processed. It is the expected output for the determination of lead. The process of removing information from an automatic device (such as a computer or sensor) and displaying it in an understandable form. A digital readout (DRO) is a numeric display, usually with an integrated keyboard and some means of numeric representation. Its integral computer reads signals generated by linear encoders or (less frequently) rotary encoders installed to track machine axes, using these measures to keep track of and display to a machine operator the work piece position (e.g..milling machines), or tool position (lathes, grinders, etc.) in space.

Digital screen to record an UV spectrograph with absorbance against the wavelengths. Definitions of recording system. Audio system for recoding sound. Type of audio system, sound system. A system of electronic equipment for recording or reproducing sound.

ADVANTAGES

- The core advantage is the accuracy of the UV-VIS spectrophotometer
- The UV-VIS spectrometer is easy to handling and use
- Provide robust operation
- UV-VIS spectroscopy is simple to operate
- Cost effective instrument
- Cover the entire of ultraviolet and visible
- It can be utilized in the qualitative and quantitative analysis
- The Derivative graph can be obtained by UV-VIS spectrophotometer
- It can be used in the degradation study of drug
- Only possible for the analytes which have a chromophore.

DISADVANTAGES

- Only those molecules are analyzed which have chromophore
- The results of the absorption can be affected by pH, temperature, contaminants, and impurities.
- Only liquid samples are possible to analyze
- It takes time to get ready to use it
- Curette handling can affect the reading of the sample

APPLICATIONS

- UV –Visible spectroscopy has many different application
- Detection of impurities

- Structural elucidation of organic compounds
- Quantitative analysis
- Qualitative analysis
- Chemical analysis
- Quantitative analysis of pharmaceutical substance
- Dissociation constant of acids and bases
- Molecular weight determination
- As HPLC detector.

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